

CHEM 116/222

Exp. 6,7: Preparing Standard Acid and Base; Using a pH Electrode for an Acid-Base Titration

Name: _____ Date: _____

TA's Name: _____

EXP. 6 DATA:**SODIUM HYDROXIDE****HYDROCHLORIC ACID**

	trial 1	trial 2	trial 3	trial 4	trial 1	trial 2	trial 3	trial 4
<i>m</i> standard (g)								
<i>V</i> titrant (mL)								
molarity								
std dev, % rel	/				/			

EXP. 7 DATA:**UNKNOWN ACID (unknown number _____)**

M base for titration		titration curve	first derivative	second derivative
mass of unknown (g)				
indicator endpt for 100 mL soln (mL)				
equivalence volume (mL)				
molecular weight (g/mol)				

UNKNOWN BASE (unknown number _____)

M acid for titration		titration curve	first derivative	second derivative
mass of unknown (g)				
indicator endpt for 100 mL soln (mL)				
equivalence volume (mL)				
molecular weight (g/mol)				

Note: include computer generated graphs of titration curve, $d(\text{pH})/dV$ and $d^2(\text{pH})/dV^2$ in report.